

Controlling reaction rate of phase transfer hydrogenation of acetophenone by application of low external electric field

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Abstract

The controllable mass transfer and reaction rate for phase transfer hydrogenation of acetophenone across a well-defined boundary were investigated. The effect of solvent was found important and 1-butanol exhibited the best performance among the five investigated homologous alcohol solvents, consistent with its higher solubility in water and greater dielectric constant. Initial reaction rates increased with increasing electric potential, consistent with enhanced mass transfer across the aqueous/organic boundary. At longer reaction times deactivation was apparent. It correlated with increasing voltage and is ascribed to lower equilibrium concentration of reactive species at the interface. External control over reaction rate was demonstrated by switching the applied electric potential over the course of the reaction. Effects of external electric field on enantioselectivity were also explored with reversal field direction. The changes correlate with catalyst decomposition.

1. Introduction

Controlling reaction rate and selectivity is a long-standing goal and challenge in chemistry.¹⁻³ Conventional ways to control chemical reactions include heating,⁴ stirring, using solvents⁵ or catalysts.⁶ Recently, inspired by the pioneering work of Shaik and co-workers,⁷⁻⁹ attention has turned to the manipulation of activation energy barriers and the stabilization of reactive species in chemical reactions by oriented external electric fields (OEEFs).^{10,11} The influence of OEEFs upon chemical reaction performance has raised a number of scientifically interesting questions, and increasing evidence has shown that OEEFs could be employed as smart reagents and catalysts in chemistry.^{10,12-15} It was proposed by Meir et al.⁸ that organic reactions can be theoretically accelerated when an external electric field is oriented along the “reaction axis” which enhances the electron flow, leading to a lower reaction barrier and the stabilization of transition states. This simulated prediction was experimentally demonstrated by Aragonés et al. in an OEEF promoted Diels-Alder reaction at the single molecule scale.¹⁶ Huang et al. also proved this idea by selective catalysis of a Diels-Alder addition reaction with an external electric field oriented along the reaction coordinate.¹⁷ More importantly, Wang et al. predicted that the reaction enantioselectivity may be controlled by aligning the electric field along the dipole moment of the reactants.² While it is potentially significant for asymmetric synthesis, there is no direct experimental evidence so far to support it.

Despite the promising prospects, there are two major challenges to be met in order to utilize OEEFs for controlling reaction rates and selectivity. One requirement is to orient the electric field along the reaction axis,¹⁰ the other is the high electric field strength (10^7 to 10^9 V/m).^{17,18} At molecular level, these challenges can be addressed by using scanning tunneling microscope break junction (STM-BJ) techniques^{16,19,20} or mechanically controllable break junction (MCBJ) techniques.^{17,21,22} Nevertheless, when conducting experiments on a bulk scale, the question of how to align the electric field along the reaction axis arises when thousands of reactant molecules are randomly oriented in the reaction system. Furthermore, applying thousands of volts to a lab-scale (or even larger scales) reactor could pose safety issues. Therefore, the significance of OEEFs on controlling reaction rates and selectivity experimentally in bulk quantity reactions for organic synthesis

remains understudied. However, apart from lowering the reaction barrier with huge energy input, a simple way to manipulate the reaction rates by external electric fields is to control the diffusion of reactive species in the reaction system,²³ and alter the local concentration of reactants.²⁴ This has been validated in our previous study that the reaction rate can be either promoted or inhibited by simply flipping the orientation of the external electric field.²⁵

To further explore the possibilities of controlling reaction rate and even selectivity with external electric field on bulk scales, catalytic transfer hydrogenation of acetophenone was studied in a biphasic liquid system. Previous studies showed that the phase transfer hydrogenation of acetophenone with sodium formate as hydrogen source is a mass transfer limited reaction happening at the organic-aqueous interface.^{25,26} When an external electric field was applied, the charged reactive species could be either transferred to the interface to promote the reaction or constrained in the corresponding liquid phases decelerating the reaction by the electrostatic forces.²⁵ This raises the possibility of manipulating the reaction rates by controlling the diffusion of reactive species in the presence of varying electric field potential for interfacial reaction/catalysis in a biphasic liquid system.

In this contribution, we have demonstrated the ability of external electric fields to control reaction rates in the catalytic phase transfer hydrogenation of acetophenone and its influence on reaction enantioselectivity. Mathematical modeling and simulation results further support the controllable migration of reactive species under external electric fields. While evincing fundamental principles, the results also show the possible decomposition effects of the external electric field on the catalyst when applied in the negative orientation, resulting in lower reaction conversion and enantioselectivity. In general, this work shows the importance of comprehensively understanding the role of external electric field in organic synthesis and catalysis on bulk scale.

2. Experimental section

2.1 Materials

All chemicals used in this work were purchased from commercial reagent suppliers and used without further purification: acetophenone (Sigma-Aldrich, 99%), sodium formate (Sigma-Aldrich, 99.0%), RuCl(p-cymene)[(S,S)-Ts-DPEN] (Sigma-Aldrich), 1-butanol (Sigma-Aldrich, 99.8%), 1-pentanol (Sigma-Aldrich, [?] 99%), 1-hexanol (Alfa Aesar, 99%), 1-heptanol (Sigma-Aldrich, 98%), 1-octanol (Sigma-Aldrich, [?] 99%).

2.2 Experimental apparatus and procedure

The phase transfer hydrogenation of acetophenone to produce 1-phenylethanol was studied in a specially made hollow cylindrical glass reactor with two ports for electrodes on the ends. The inner diameter of the reactor is about 20 mm. Details of the experimental unit may be found elsewhere.²⁵ Briefly, the biphasic system was set up with a sodium formate/water solution as the more dense aqueous phase and acetophenone/organic solvents as the less dense organic phase in which the Ru complex (RuCl(p-cymene)[(S,S)-Ts-DPEN]) was dissolved. The organic solvents used in conjunction with acetophenone were 1-butanol, 1-pentanol, 1-hexanol, 1-heptanol, and 1-octanol. An external electric field (DC voltage) from a high voltage power supply (PASCO SF-9585A) was applied across the two phases without agitation. Stainless steel or titanium electrodes were used. In a single experimental study, 0.34 g of sodium formate was fully dissolved in 1.5 mL of deionized water, while 117 μ L of acetophenone was mixed with 1.5 mL of organic solvent to dissolve 0.0064 g Ru catalyst. The reactions were conducted at room temperature and pressure with positive or negative electric field applied during reaction. The detailed orientations of the electric field have been described previously.²⁵ A positive orientation in this work is defined as the direction of the electric field when the anode is placed in the organic phase with the cathode in the aqueous phase. The conversion of acetophenone, total yield of 1-phenylethanol, and enantioselectivity of (S)-1-phenylethanol were determined by gas chromatography analysis.

2.3 Gas chromatography (GC) analysis

10 μL of liquid samples from the organic phase were withdrawn after each reaction and diluted to 1.5 mL with methanol for gas chromatography analysis. An Agilent 6890N gas chromatograph equipped with a CP-Chirasil-Dex CB column (25m \times 0.25 mm) and a flame ionization detector was used to identify reaction products and determine enantioselectivity under the following conditions: helium as carrier gas (3 mL/min), inlet and detector temperature 250 $^{\circ}\text{C}$, oven temperature 115 $^{\circ}\text{C}$.

2.4 Nuclear Magnetic Resonance (NMR) spectroscopy analysis

Four samples were prepared for NMR analysis. Sample 1 was prepared by dissolving 5.8 mg of $\text{RuCl}(\text{p-cymene})[(\text{S,S})\text{-Ts-DPEN}]$ in 500 μL of MeOH-d_4 (Cambridge Isotopes). Sample 2 was obtained by dissolving 6.4 mg of $\text{RuCl}(\text{p-cymene})[(\text{S,S})\text{-Ts-DPEN}]$ in a mixture of 1.5 mL 1-butanol and 117 μL acetophenone. Samples 3 and 4 were collected from reactions under positive and negative 15 V of electric potential respectively after 24 h with titanium electrodes and 1-butanol as solvent. 500 μL of each liquid sample were pipetted into a 5 mm NMR tube (Wilmad Lab-Glass, Vineland, NJ). All NMR spectra were acquired on a 500 MHz Bruker AVIII spectrometer equipped with a cryogenically cooled X-channel observe probe. All data was analyzed using MestreNova NMR software (Santiago De Compostela, Spain).

2.5 X-ray florescence (XRF) analysis

A handheld XRF analyzer (Olympus, Delta Professional) with Ru anode and silicon drift detector was used for fast determination of the metal contents in the degradation products of stainless-steel electrode. These solid samples were collected from the interface after reaction and dried prior for analysis. Precious metal mode with a targeted collimator was employed during analysis. All samples were scanned three times and analyzed with the workstation setup.

3. Theoretical modeling and simulation

The transport of the charged reactive species, referred to as electromigration in the present study, was modeled based on the Nernst-Planck equation.²⁷⁻²⁹ The molar flux of species i , J_i , is expressed in Equation (1), describing the diffusion by concentration difference and migration by electrostatic force. The convection term is not included due to the fact that the experiments were conducted in a quiescent hydrodynamic environment.

$$J_i = -D_i \nabla c_i - \frac{D_i z_i F}{RT} c_i \nabla \Phi \quad (1)$$

where D_i is the diffusion coefficient (m^2/s), c_i is the ion concentration (mol/m^3), z_i is the valence number, F is the Faraday constant, R is the gas constant, T is the absolute temperature (K), and Φ is the electric potential (V). The Wilke-Chang equation was applied to estimate the mutual diffusion coefficient if necessary.³⁰

The electric field in the domain is described by the Poisson differential equation³⁰:

$$\nabla \bullet (\epsilon_0 \epsilon_r \nabla \Phi) = -\rho_v \quad (2)$$

where ϵ_0 is the dielectric constant of the free space, ϵ_r is the relative permittivity of materials, ρ is the electric space charge density.

The equations were solved using finite element method in COMSOL Multiphysics 5.4. Simulation was conducted to understand the mass transfer of reactive species in the proposed reaction system, assuming no reaction nor flux at/across the interface. Concentration profiles of Ru catalyst and formate at the interface were calculated under various external electric fields.

4. Results and discussion

4.1 Solvent effects

The effects of the different organic solvents used in conjunction with the acetophenone reactant were first studied as their different physical and chemical properties could significantly affect the reaction performance, especially when an external electric field is applied. Five primary alcohols, namely 1-butanol, 1-pentanol, 1-hexanol, 1-heptanol, and 1-octanol as organic solvents were investigated due to their ability to maintain a

stable biphasic reaction system. The experimental results are presented in Figure 1, showing total conversion values and product yields. Table 1 lists the physical and chemical properties of the five solvents. No corresponding aldehyde by-products were detected by GC analysis after the reactions, which confirms the role of sodium formate as hydrogen source and the concept of phase transfer hydrogenation. As we can see, addition of 1-butanol resulted in the best performance on promoting the transfer hydrogenation of acetophenone with conversion of 84.7% and yield of 76.8% after 24 h of reaction under 15 V of positive DC voltage with stainless-steel electrodes. The conversion and yield decreased gradually with the increase of the molecular weight of these homologous alcohol solvents.

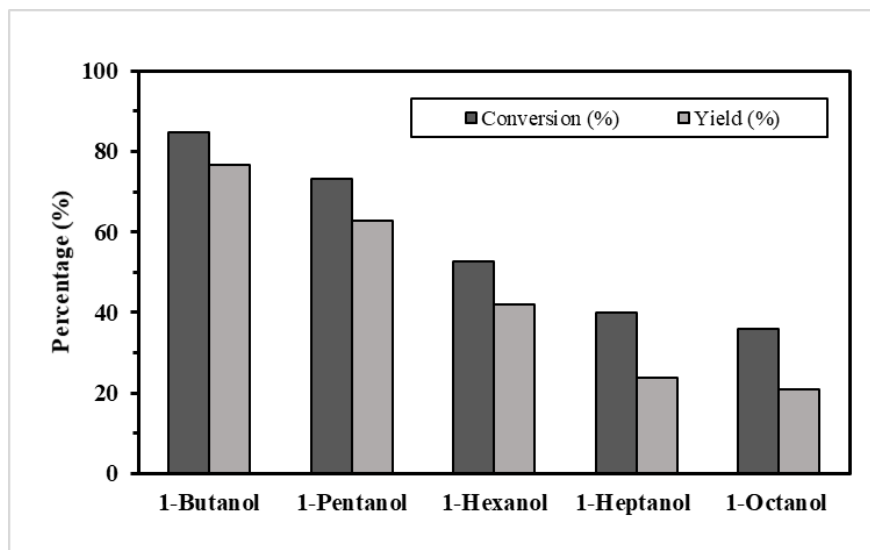


Figure 1. Acetophenone conversion (mol%) and 1-phenylethanol yields (mol%) under 15V of positive external electric field for 24 h with different organic solvents and stainless-steel electrodes.

One possible reason for the significantly different performance may be their respective solubilities in water. Compared to the other solvents, 1-butanol has the highest solubility in water, which initially promotes most of the interaction between the reactants at the interface, resulting in the highest conversion and reaction rate (more than 2 fold faster than that with 1-octanol as solvent). On the other hand, these alcohols are protic solvents, which possess the ability to generate protons by dissociation. With the largest dielectric constant among the studied solvents, 1-butanol has the greatest potential to polarize and generate protons under an external electric field. These protons could further promote the dissociation of the neutral Ru compound ($\text{RuCl}(\text{p-cymene})[(\text{S-S})\text{-Ts-DPEN}]$) into the cationic Ru complex (species **2** in Scheme 1) by removing the chloride ion,^{31,32} therefore accelerating and promoting the generation of the catalysts and ultimately increasing the production rate. While the dipole moment and electrical conductivity might be important in electrostatically promoted reactions according to our previous investigation²⁵ and Shaik and coworkers,² the differences of these properties between the five solvents are small. Accordingly, it is difficult to attribute the differences in reaction performance to dipole moment and conductivity in the present study.

Table 1. Properties of the studied organic solvents.

Solvent	Electrical conductivity (ns/cm)	Solubility in water (25°C/mass%) ³³	Dipole(D) ³³	Dielectric constant (20°
1-butanol	36 ^{34,35}	7.3	1.66	17.84
1-pentanol	41 ^{34,35}	2.14	1.7	15.13 ^a
1-hexanol	—	0.59	—	13.03
1-heptanol	—	0.164	—	11.75

Solvent	Electrical conductivity (ns/cm)	Solubility in water (25°C/mass%) ³³	Dipole(D) ³³	Dielectric constant (20°
1-octanol	140 ³⁶	0.046	1.76	10.3

a: This value was obtained at 25°C.

4.2 Electric field potential

Various values of positive external electric field potential were applied to determine its effects on mass transfer and kinetics in the proposed reaction system. The conversion of acetophenone under different positive external electric field strengths, recorded at different times are shown in Figure 2. The data in this Figure are only for the acetophenone/1-butanol mixture using stainless steel electrodes. It was observed that the initial reaction rates increased with the increase of the electric potential. The conversion of acetophenone in the presence of an applied voltage of 100 V was observed to reach 44.5% after a reaction time of 2 hours, which is nearly 4 fold of the conversion attained at 15 V. However, at longer reaction times, the rate of increase was observed to decline. The decline was greater at the higher voltages. For example, after 8 h, a higher conversion was achieved at 15 V compared to the values observed at 30 V and 50 V. The total conversion at 15 V reached to 84.7% after 24 h which is about 20% higher than that at 30 V and 40% higher than that at 50 V.

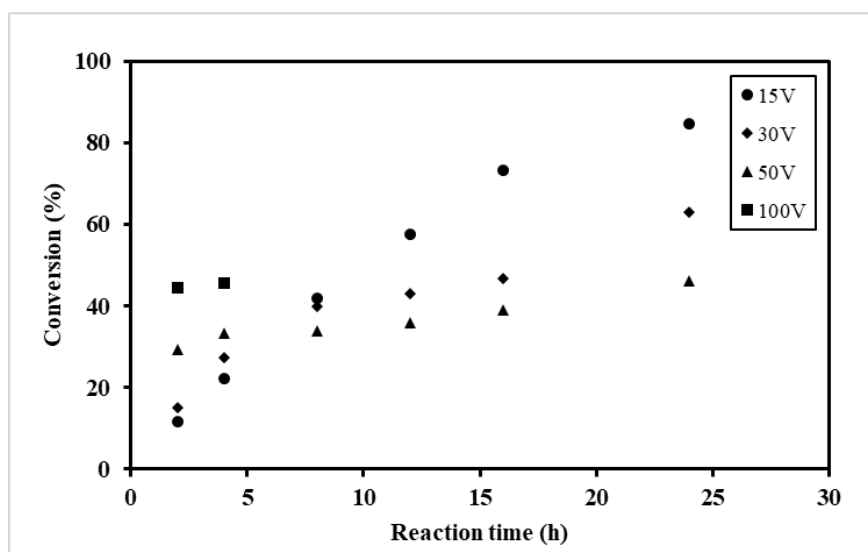


Figure 2. Acetophenone conversion (mol%) under different positive external electric field for 24 h with stainless-steel electrodes and 1-butanol as solvent.

The increase in conversion with increased electric potential in the first few hours is consistent with the role of external electric field in controlling the reaction rates of the phase transfer hydrogenation reaction by controlling mass transfer of the charged reactive species. This is further supported by the simulation results in Figure 3. As we can see, the rates of electromigration of both formate and Ru catalyst were increased with the increase of voltage at the beginning. The enhancement of reaction rate observed at higher electric field is consistent with accelerated rates of transport of both formate anions and the cationic Ru catalysts to the liquid-liquid interface, resulting in more rapid reaction compared to that under lower electric potential. The simulation results also suggested that the equilibrium concentration of Ru catalyst at the interface is lower at higher voltage, which explains the corresponding lower conversion at longer reaction times. Additionally, a negatively charged alkoxide intermediate could be formed during the reaction as suggested by Pavlova et al.³⁷ As the reaction continues, a portion of these alkoxide anions could be attracted to the anode electrode and constrained in the region close to it, preventing the reactive species from interacting and retarding the

reaction rates. This inhibition is greater when a greater electrostatic force is applied, leading to the lower reaction rate as observed at an elevated electric potential for longer reaction times.

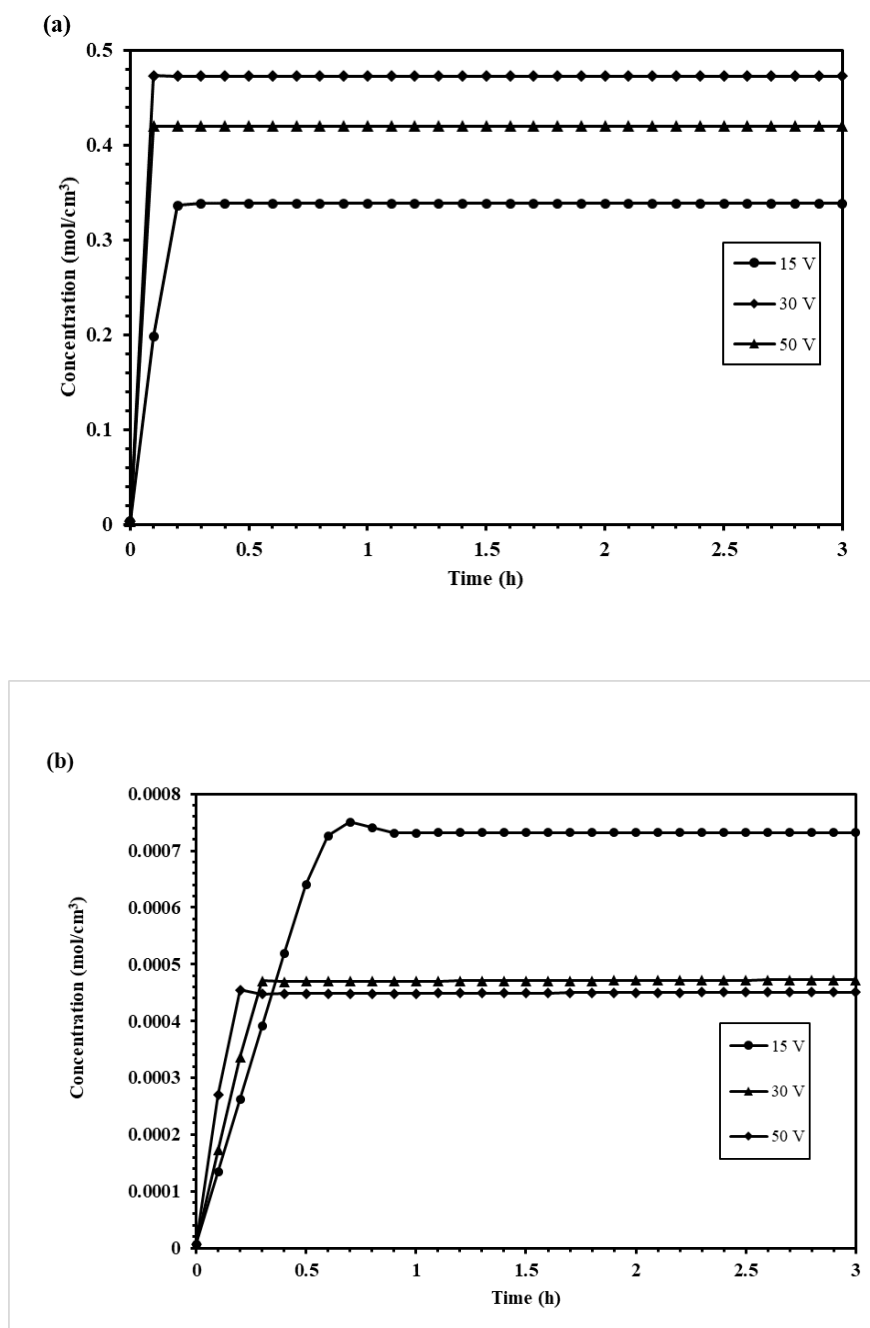


Figure 3. Simulated concentration profiles at the interface under various electric fields. (a) Concentration profile of formate ions, (b) concentration profile of Ru catalyst.

On the other hand, an external electric field induced current flow in the reaction system along with degradation of the anode electrode (immersed in the organic phase) were observed. Additionally, gas bubbles

of unknown composition were generated on the cathode. The degradation of the electrode was confirmed by weight loss measurement (see Cases 1, 2 and 3 in Table 2). The degradation products which formed a thin layer of black slurry at the liquid-liquid interface were characterized by XRF analysis. The analysis showed that the slurry consisted mainly of Fe, Ni, and Cr with a small amount of Ru. For reactions under higher electric potentials, higher current flows and greater degradation rates of the electrode were observed as shown in Table 2 (Cases 1, 2, and 3). The decomposition of the electrode is consistent with galvanic corrosion, in which hydrogen is produced at the cathode with the corrosion occurring at the anode. At higher electric potentials, stronger galvanic corrosion effects would be expected, leading to greater electrode degradation. With more metal ions released from the degradation, higher current flows would be observed due to increased conductivity of the liquid phase. The degradation products formed at the interface could also contribute to the deceleration of the reaction by inhibiting mass transfer to the catalyst at the interface. At higher electric potential, a higher degradation was observed, leading to a higher blocking coverage of the interface and finally achieving a lower reaction rate and conversion.

To eliminate the negative effects of electrode degradation and clarify the inhibiting effects of high electric potential, chemically stable titanium electrodes were employed, and the results are recorded as Cases 4, 5, and 6 in Table 2, compared to those with stainless steel electrodes (Cases 1, 2, and 3). For experiments with different electric potential applied, the average degradation rate of titanium electrodes is less than 0.1% w/w, and no apparent black slurry was observed. Although there was negligible evidence of electrode degradation of titanium electrodes, reaction performance showed the same trend as observed with stainless steel electrodes, when the applied electric potential was increased. This suggests that the actual electrode degradation plays a minimal role in changes in reaction performance, thus strongly suggesting the controllable migration of Ru catalyst and alkoxide ions under electrostatic force as being the main reason for the reaction deactivation and the importance of external electric field in controlling reaction rates. Furthermore, in spite of apparent reaction deactivation, a constant product ((S)-1-phenylethanol) enantioselectivity was achieved. This is further evidence that the deactivation is most likely due to the controllable migration of reactive species rather than catalyst decomposition.

A further set of experiments was conducted to investigate the reaction performance when the applied electrical potential was varied over the course of the reaction. With reference to Table 2, the data shown in the last three rows list the results of these experiments, labelled Cases 7, 8, and 9. In case 7, after reacting for 8 h under 30 V, the electric potential was switched to 15 V for 16 h, and the reaction yield reached to 67.5%, which is an increase of 7.4% compared to that with solely 30 V applied for 24 h (Case 5). In case 8 and 9, 30 V or 50 V were applied for 4 h respectively followed by 15 V applied for 20 h. In these two cases, the yields of 1-phenylethanol increased when comparing to those with solely 30 V or 50 V applied for 24 h. These results suggest that the reaction could be externally controlled by simply switching the applied electric potential over the course of the reaction.

Table 2. Summary for experiments under various positive external electric fields.

Case	Voltage (V)	Time (h)	Electrode	Current (mA)	Weight loss (% w/w)	Yield (%)	ee (%)
0	0	24	—	N/A	~0	49.8	94.6
1	15	24	SS	1.2	3.5	76.8	93.5
2	30	24	SS	1.7	4.3	55.5	93.4
3	50	24	SS	2.6	4.4	32.6	93.3
4	15	24	Ti	0.2	0.085	67.4	94.5
5	30	24	Ti	0.5	0.085	60.1	94.5
6	50	24	Ti	1.6	0.09	26.6	94.5
7	30-8h + 15-16h	30-8h + 15-16h	Ti	0.2	0.09	67.5	94.5
8	30-4h + 15-20h	30-4h + 15-20h	Ti	0.2	0.02	63.9	94.3
9	50-4h + 15-20h	50-4h + 15-20h	Ti	0.2	0.1	47.1	94.3

4.3 Proposed mechanism

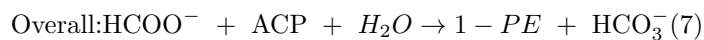
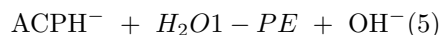
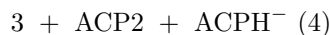
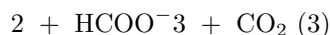
Based on the observations, a stepwise mechanism for the phase transfer hydrogenation of acetophenone with aqueous formate as hydrogen source and Ru-TsDPEN catalyst under an external electric field is proposed in Scheme 1. The cationic Ru complex **2** is generated by removal of chloride after compound **1** dissolves in 1-butanol. Following the formation, species **2** undergoes electromigration to the interface where a hydride is transferred to **2** from a formate to generate catalyst **3**, with the release of CO₂.²⁶ Instead of transferring both hydrogen to the substrate concertedly,³¹ only the hydride on Ru-H in catalyst **3** is transferred to acetophenone with the proton on the NH₂ moiety retained,³⁷ resulting in the regeneration of species **2** and the formation of the corresponding alkoxide intermediate due to the stabilization of hydrogen bonding by water.³⁷ Finally, the alkoxide picks up a proton from water to form the product,³⁷ and the generated hydroxide further reacts with CO₂ to form bicarbonate as observed in our previous study.²⁵

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image5.emf available at <https://authorea.com/users/332910/articles/459244-controlling-reaction-rate-of-phase-transfer-hydrogenation-of-acetophenone-by-application-of-low-external-electric-field>

Scheme 1. Proposed mechanism for the phase transfer hydrogenation of acetophenone with aqueous formate as hydrogen source and Ru-TsDPEN catalyst.

The stoichiometry and rate expression consistent with the proposed mechanism are summarized in Equations (3) - (9), assuming that the coordination of formate and decarboxylation are reversible and equilibrated prior to hydrogen transfer,²⁶ where K_1 is the equilibrium constant, and k_1, k_2, k_3 are the respective reaction rate constants. *ACP* and *ACPH*⁻ represent acetophenone and its corresponding alkoxide, while 1-*PE* refers to 1-phenylethanol and $[Ru]$ represents the total concentration of catalyst. As CO₂ rapidly reacts with hydroxide during the reaction, the concentration of CO₂ is expected to be small, therefore the rate expression can be further simplified as a second order kinetics over the concentration of catalyst and substrate (Equation (9)). This is consistent with the study by Wu et al.²⁶ and with the above observation showing that the reaction rates were limited by the concentration of Ru catalyst and the alkoxide at the interface region. Both of these concentrations were affected by the electromigration under the influence of external electric fields. These observations suggest an important impact of electric fields on kinetics of interfacial reactions involving reactive ions.



$$\text{Rate expression: } r = \frac{K_1 k_1 [Ru][ACP][\text{HCOO}^-]}{[\text{CO}_2] + K_1 [\text{HCOO}^-]} \quad (8)$$

$$\text{When } [\text{CO}_2] \ll K_1 [\text{HCOO}^-], r = k_1 [Ru][ACP] \quad (9)$$

4.4 Electric field polarity and product enantioselectivity

As shown in Table 2, the enantioselectivity of (S)-1-phenylethanol remained constant with respect to electric field potential applied in the positive orientation. Since the orientation of OEEF was proposed as being important for controlling reaction selectivity, a reversed external electric field was then applied to the reaction system. The results are presented in Table 3. In these cases, the reaction was inhibited when the electric potential was applied in the negative orientation, which further confirms the hypothesis of controllable mass transfer and reaction rates by external electric fields, in accord with past findings. The yield of 1-phenylethanol increased with the increased negative electric field strength, which was unexpected as a larger

inhibiting electrostatic force would be expected to result in a lower reaction rate and conversion. This increase is likely due to the significant increase on the current flow (Figure 4) in the reaction system. Increase in current flow is consistent with the local temperature increase due to joule heating, which would compete with the inhibiting effects of higher negative electrostatic force and possibly promote the reaction.

Table 3. Summary for experiments with various negative external electric fields applied.

Voltage (V)	Voltage (V)	Time (h)	Electrode	Current (mA)	Yield (%)	EE (%)
0	0	24	—	N/A	49.8	94.6
-5	24	24	Ti	—	16.9	91.0
-10	24	24	Ti	—	22.4	86.0
-15	24	24	Ti	—	23.0	66.1
-20	24	24	Ti	—	45.6	82.9
-30	24	24	Ti	—	48.4	75.2

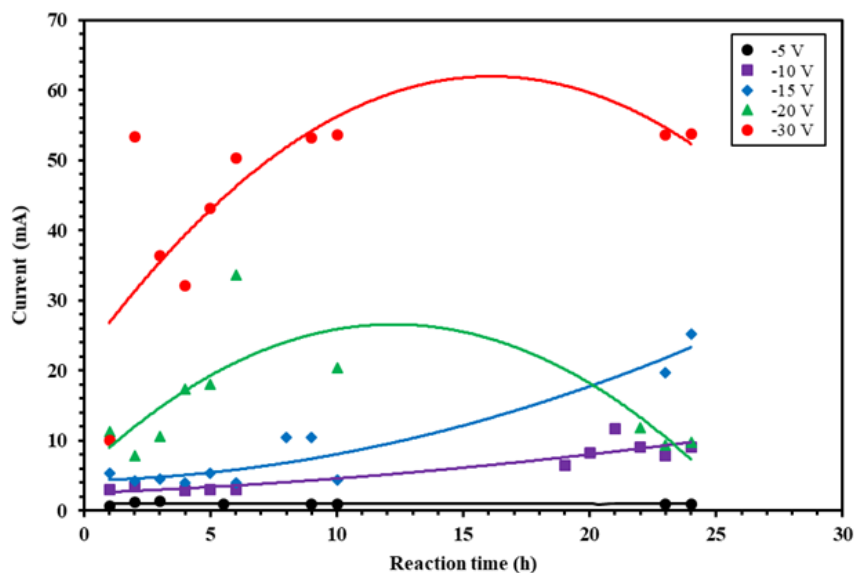


Figure 4. Current flow at different reaction time under various negative external electric fields.

On the other hand, with the increase of the inhibiting electric potential, decreased (S)-1-phenylethanol enantioselectivities were observed. However, these changes in the enantioselectivity are probably not due to the ability of OEEF to control reaction selectivity but are more likely attributed to the catalyst decomposition. Figure 5 showed the NMR analysis for the identification of the structures of the catalysts before and after reactions. Samples 1 (spectrum A) and 2 (spectrum B) were obtained by dissolving the Ru catalysts in CD₃OD and reaction mixture (1-butanol and acetophenone mixture) respectively, and peaks from the Ru catalyst are clearly present at approximately 6.57 and 6.79 ppm. For sample 3 (spectra C), collected from the reaction run at positive 15 V for 24 h, small signals (due to the low initial concentration and difficulty of detection in ¹H NMR spectra) for the Ru catalyst can be seen. However, the NMR spectra for sample 4, from the reaction run at negative 15 V for 24 h, showed no signals for the Ru catalyst. It remains unknown if the decomposition of the Ru catalyst was caused by the negative external electric field, and further investigation is required.

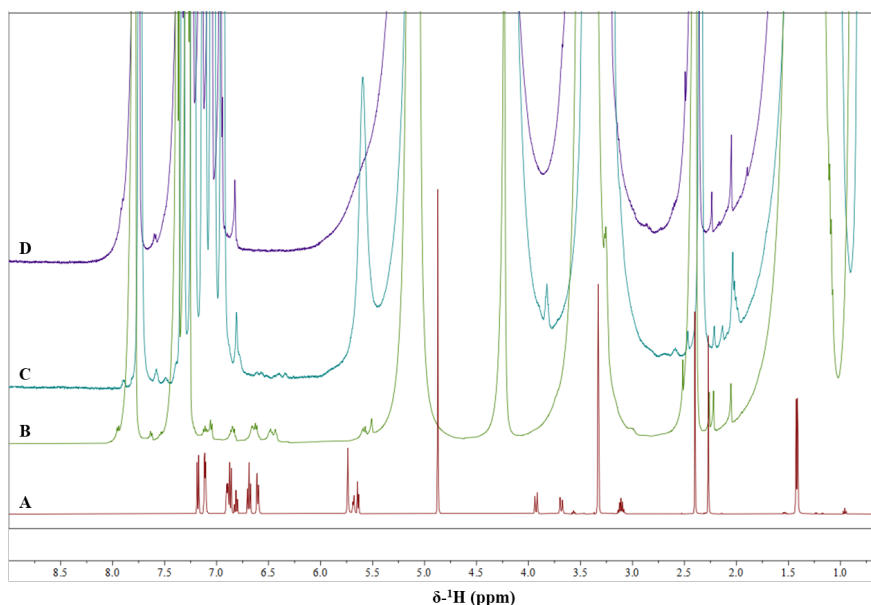


Figure 5. Stack plot of 1D ^1H NMR spectra of samples 1 (A), 2 (B), 3 (C) and 4 (D). To highlight the minor species, spectra B was expanded vertically by 8 times, and spectra C and D were expanded by 24 times.

5. Conclusion

The reported work demonstrated experimentally controllable reaction rates by low voltage external electric fields for the catalytic transfer hydrogenation of acetophenone. The influence of applied electric potential on reaction performance was evaluated experimentally and the controllable mass transfer of reactive species by electric fields was supported by simulation results. Significant improvements were shown at low values of voltage relative to a control of zero volts, with an optimum voltage of positive 15 V. Increases beyond this value to positive 30 V and 50 V showed significant reduction in performance due to lower equilibrium concentration of Ru catalyst at the interface and possible formation and migration of the alkoxide intermediate during reaction. The degradation products from the electrolytic corrosion of the stainless-steel electrodes did not significantly impact the hydrogenation reaction. The influence of different external electric fields on the concentrations of reactive species at the interface was suggested as the main reason for the differences in reaction performance. Enantiomeric excess values were measured, and no significant changes were observed with variation in the positive voltage values. A mechanism and reaction rate expression were proposed based on the observations, suggesting the importance of external electric fields in reactions involving reactive ions.

The direction of the applied electric field was proved to be important as both reaction conversion and enantioselectivity were significantly reduced when the electric field was in the negative orientation. A monotonic increase in conversion was observed as voltage increased from negative 5 V to negative 50 V due to possible temperature increase resulted from the increase of current. Catalyst decomposition by the influence of negative voltages was concluded as the main reason for reduced enantioselectivity, as implied from the NMR analysis.

Overall, this work extends the scope for controllable synthesis of organic reactions conducted in two phase liquid systems and raises an opportunity for green engineering with minimal energy consumption in the proposed reaction system. It is also important for understanding the influence of external electric fields on mass transfer and kinetics in biphasic liquid systems used for organic synthesis and interfacial catalysis on macroscale.

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